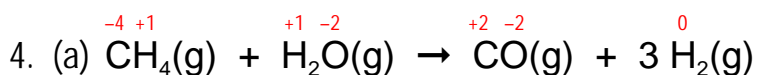
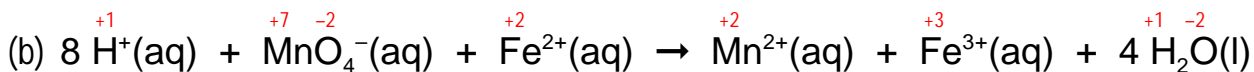


## ANSWERS

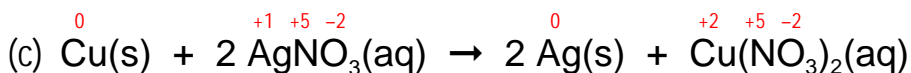
3. (a) 0 (element) (f) -3  
 (b) +6 (g) +3  
 (c) -2 (h) +7  
 (d) -1 (monatomic ion) (i) -2  
 (e) +2 (j) +4



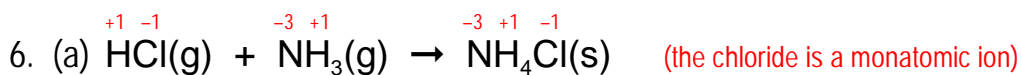
The carbon is oxidized and the hydrogen is reduced.



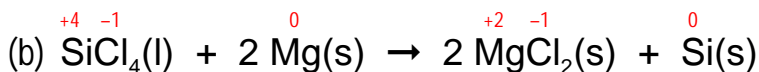
The iron is oxidized and the manganese is reduced.



The copper is oxidized and the silver is reduced.



This is NOT a redox reaction (no changes in oxidation numbers).

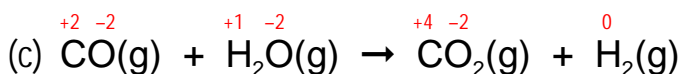


SiCl<sub>4</sub> is the oxidizing agent.

Mg is the reducing agent.

Magnesium is oxidized.

Silicon is reduced.



H<sub>2</sub>O is the oxidizing agent.

CO is the reducing agent.

Carbon is oxidized.

Hydrogen is reduced.